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Chemistry lab manual class 11 cbse pdf

Blueprint's Chemistry Lab Manual for class 11 covers the entire syllabus strictly in accordance with the guidelines prescribed by CBSE. The manual is especially designed and helps the learners to develop skills in scientific enquiry and laboratory techniques and use them to interpret the results of experiments. The manual strives to deliver: Development of a strong conceptual base. Familiarity with the required apparatus and the properties of various elements. Helps in qualitative analysis and quantitative estimation. Ability to predict outcomes on the basis of observations made. Complete knowledge of precautions which are to be taken during experiments. Facilitating interactive analysis through viva-voce questions which help to understand the laws and principles related to the experiments. Necessary additional information through appendix. (List of Chemicals and apparatus, log tables etc.) ADDITIONAL FEATURE Lab Experiment with QR Code (Scan to Experience Visualisation) ADDITIONAL INFORMATION ISBN 978-93-85470-41-7 AUTHOR Anamika Saxena EDITION Latest Edition PAPERBACK/HARD BOUND Hard Bound Customer Reviews Add your review: Showing 1 to 13 of 13 (1 Pages) Class 11 Chemistry Lab manual prepared in accordance to the new syllabus for Clas 11 chemistry prescribed by CBSE. Our expert teachers have taken utmost care while preparing Class 11 chemistry lab manual More Resources for CBSE Class 11 Price - 76 152 Price - 394 438 Price - 367 408 Price - 376 418 Price - 270 318 Price - 125 249 Price - 377 419 Price - 332 369 Price - 377 419 Price - 356 419 Price - 365 429 Price - 382 449 Price - 382 449 Price - 339 399 Price - 287 319 Price - 287 319 Price - 260 289 Price - 314 369 Price - 571 688 Price - 1243 1516 Price - 1632 2015 Price - 1624 2005 Price - 108 120 Price - 209 232 Price - 224 249 Price - 134 149 Price - 424 499 THIS LAB MANUAL IS IN ACCORDANCE WITH THE LATEST SYLLABUS. THE EXPERIMENT PART IS FOLLOWED WITH BLANK PAGES FOR EXPERIMENT WRITING. INTRODUCTION L ABORATORY work has special importance in the learning of science as scientific principles develop and grow on the basis of laboratory work. Chemistry is an experimental science; the concepts learned in the theory classes are better understood through experimentation. Laboratory work provides an opportunity to observe many of the chemical phenomena under controlled laboratory conditions and workout a problem through the method of inquiry. In other words, it provides you with ample opportunity to become a keen observer and to draw inferences and explain results. The training in laboratory work helps to develop skills for handling apparatus and equipment and carry out experiments. In this way, the experimental work helps to promote scientific temper and adopt a cooperative attitude. Working in the laboratory provides a platform for trying novel and creative ideas and giving them concrete shape. Before you become familiar with the scientific procedures and experimental skills and start working in the laboratory, you should be well acquainted with the chemistry laboratory. You should notice the facilities provided in the laboratory and on your working table. You will notice that your table is provided with a water-tap, gas-tap, Bunsen burner spirit lamp/kerosene lamp, a reagent shelf and a bin for disposing waste material. You will find that some reagents are kept on the shelf fixed on the table while some reagents are kept on the shelf fixed on sidewalls. Reagents kept on the shelf of the table are frequently required while those on the shelf fixed on walls are less commonly used. Besides the facilities on the table, you will see that there are exhaust fans fitted on the upper portion of the wall opposite to the doors and windows and are placed close to the level of the ceiling. This facilitates the exhaustion of harmful fumes and circulation of fresh air in the laboratory. For this purpose, there are also enough number of windows in the laboratory. Keep these open while working. A fume cupboard is provided in the laboratory for performing those experiments in which fumes are produced. It is strongly advised that you should become fully familiar with your chemistry laboratory, laboratory practices and procedures, and more importantly the precautions to be taken while working in the laboratory. The environment in the chemistry laboratory is somewhat special in the sense that it can be a place of joy, discovery and learning. It can also become a place of frustration and danger. Frustration – if you come unprepared and neglect recording the important data properly, and danger – if you do not follow the precautions properly in conducting the experiments where potential danger exists. Click on the below link to download NCERT Class 11 Chemistry Lab Manual Introduction Click to View or Download pdf file > Click for more Chemistry Study Material > Latest Syllabus for Class 12 for 2021 2022 Latest Syllabus for Class 11 for 2021 2022 Latest Syllabus for Class 10 for 2021 2022 Latest Syllabus for Class 9 for 2021 2022 CBSE has issued the latest syllabus for the academic year 2021 2022 which is applicable for all... The Central Board of Secondary Education (CBSE) today, announced a suggested competency-based assessment framework to strengthen India's existing school education system for secondary level (classes 6-10) and improve the overall learning outcomes of students across... This has a reference to the letter no CBSE/COORD/2021 dated 11.02.2021 regarding conduct of the School-Based Assessment (Internal Assessment/Practical/ Project) of Class-XII and uploading the marks by 11.06.2021. It has been observed that some schools have not been... ON POLICY Question.1 How CBSE will declare the result of Class-X? Answer. Results of Class X Board will be declared on the basis of an objective criterion developed by the Board vide Notification no.CBSE/CE/2021 dated 01.05.2021. Question.2. If any candidate is not... Looking to the present situation of the pandemic and school closures, and also taking in account the safety and well-being of the students, it is decided as follows: 1. The Board Exams for Class XII to be held from May 4th to June, 14th, 2021 are hereby postponed.... There are various experiments for CBSE class 11 chemistry students, who are going to attempt their practical exams soon. Students can also find the entire CBSE practical syllabus of Class 11 chemistry, as given below. The manual of Chemistry Lab of Class 11 prepared in accordance with the latest syllabus for Class 11 chemistry prescribed by the board of CBSE. At Vedantu, our expert and teachers experts have taken the utmost care while preparing for the lab manual of Class 11 chemistry. Let us look at how to write chemistry practical class 11 CBSE. 1. Basic Laboratory Techniques Cutting glass tube and glass rod Drawing out a glass jet Bending a glass tube Boring a cork 2. Characterization & Purification of Chemical Substance Boiling point determination of an organic compound. Melting point determination of an organic compound. Crystallization involving an impure sample of any one of the following: Alum, Benzoic Acid, Copper Sulphate. 3. Experiments Related to pH Change. Any one of the following experiments: Determination of pH value of a few solutions obtained from the fruit juices, solutions of both known and varied concentrations of acids, bases & salts using the universal indicator or pH paper. Comparing the pH value of the solutions of both strong and weak acid of similar concentration. Study the pH value change in the titration of a strong acid with a strong base using the universal indicator. b. Study of pH value change by the common-ion effect in the case of both weak acids and weak bases. 4. Chemical Equilibrium Any of the experiments from the following. Study on the shift in equilibrium between the thiocyanate ions and ferric ions by increasing or decreasing concentration of any of the ions. Study on the shift in equilibrium between the chloride ions and [Co(H₂O)₆]²⁺ by concentration change in either of the ions. 5. Quantitative Estimation Preparing the standard solution of oxalic acid. Using a chemical balance. Preparation of standard solution of the sodium carbonate. Strength determination of a given solution of sodium hydroxide by titrating it against oxalic acid's standard solution. Strength determination of a solution of hydrochloric acid that is given by titrating it against the standard sodium carbonate solution. 6. Qualitative Analysis 1. Determination of one anion and cation in salt that is given. NH₄⁺, Pb²⁺, As³⁺, Cu²⁺, Al³⁺, Mn²⁺, Fe³⁺, Ni²⁺, Co²⁺, Zn²⁺, Ca²⁺, Ba²⁺, Sr²⁺, Mg²⁺+S²⁻, CO₃²⁻, SO₃²⁻, NO₃⁻, NO₂⁻, Cl⁻, I⁻, C₂O₄²⁻, PO₄³⁻, CH₃COO⁻ (Note: Here, we have excluded the insoluble salts) 7. Detection of Sulphur, Nitrogen, Chlorine in the Organic Compounds Project The scientific investigations include laboratory testing, collecting information from various sources. A few suggested projects are: Study of purification of water methods. Checking drinking water bacterial contamination by sulphide ions testing. Testing the presence of iron, hardness, chloride, fluoride, and more related terms depending upon the regional difference in the drinking water and the study of causes of the presences of these ions above the given or permissible limit (if any). Acidity study of various samples of tea leaves. Study of the effect of acids & bases on the tensile strength of fibers. Determination of the evaporation rate of various liquids. Analysis of vegetable & fruit juices for their acidity. Important Note: Any other investigatory project that involves about 10 periods of work is chosen with the teacher's approval. Practical Examination for the Visually Impaired Students of Class XI Note: The same guidelines and evaluation scheme for the visually impaired students as given for Class XII can be followed. 1. List of the Apparatus to Identify for the Assessment in Practicals (for all the experiments) Tripod stand, Beaker, Wire gauze, funnel, glass rod, filter paper, test tube, Bunsen burner, test tube stand, test tube holder, dropper, ignition tube, tongs, china dish, tripod stand, funnel, wire gauze, standard flask, Bunsen burner, pipette, conical flask, burette, funnel, dropper, clamp stand, filter paper, and wash bottle. Procedure or Setup of the apparatus Odor detection in the qualitative analysis 2. List of the Experiments Characterization & Purification of the Chemical Substances Crystallization of an impure sample of any of copper sulphate, and benzoic acid. Experiments based on the pH Determination of pH value of some solutions that are obtained from the fruit juices, solutions of the known and varied acid, bases, and salt concentrations using the pH paper Comparison of the pH value of the solutions of both strong and weak acids having the same concentration 3. Chemical Equilibrium Study the equilibrium shift between the ferric ions and the thiocyanate ions by either increasing or decreasing their concentration. Study about the equilibrium shift in the [Co(H₂O)₆]²⁺ and the chloride ions by changing either of the ions' concentration 4. Quantitative Estimation Preparing the standard solution of oxalic acid. Determination of a given sodium hydroxide solution's molarity by titrating it against the standard solution of oxalic acid 5. Qualitative Analysis Determination of one anion and cation in a given salt Cations - [NH₄]⁺ and Anions - [CO₃]²⁻, [SO₃]²⁻, S²⁻, CH₃COO⁻, Cl⁻ (Note: Here, we have excluded the insoluble salts) Halogen Detection in the organic compound that is given. Nitrogen Detection in the organic compound that is given.

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